

The **pH scale: Basics** simulation explores the pH of acids and bases in everyday life, and how pH is affected by dilution with water.

Insights into Student Use

After using indicators like litmus or pH paper, students may think that the color of the liquid is related to pH. To tackle this idea, we show battery acid and drain cleaner with the exact same color.

Model Simplifications

pH of everyday liquids: For liquids with a range of measured pH values, an average value from the literature was used.

Dilution: The simulation does not account for different acid dissociation constants (K_a) for each liquid when calculating pH after dilution. We make the simplification that any increase in the concentration of the major ion is due to the ions already present in the added water. For example, if students add 100 mL of water to an acidic solution, then the number of moles of H_3O^+ increases by 1×10^{-8} mol. The concentration of the minor ion is then calculated using the self-ionization constant for water (K_w). These calculations account for the leveling effect of water.

Suggestions for Use

Sample Challenge Prompts

- Classify solutions as acids or bases, given their pH.
- Predict if the pH of your solution will increase or decrease after you add water.
- Describe two different ways you could fill the beaker with a solution with pH 6.00. Is it possible to use hand soap to do this? Explain.

See all published activities for pH Scale: Basics [here](#).

For more tips on using PhET sims with your students, see [Tips for Using PhET](#).

Carpenter, May 2015