

Reactions and Rates 4
Also uses **Salts & Solubility** and
States of Matter

Clicker Questions
LeChatlier's Principle

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Learning Goals

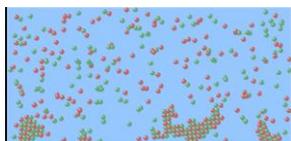
Students will be able to:

- Explain how to make equilibrium systems change and predict what changes will happen.
- Compare and contrast salt-solution, phase, and chemical equilibriums.

If you add water to this salt solution, what will happen?



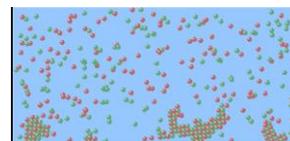
- The system will shift to the right
- The system will shift to the left
- LeChatlier's principle doesn't apply to physical systems



If you increased the air pressure above this salt solution, what will happen?



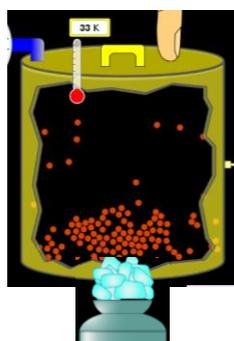
- The system will shift to the right
- The system will shift to the left
- This system would not be effected by pressure changes.



If you cooled the container, what will happen?



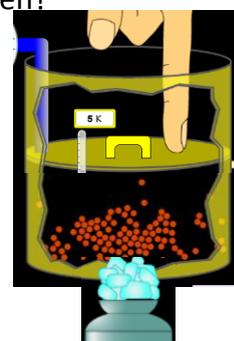
- The system will shift to the right
- The system will shift to the left
- This system is not effected by temperature



If you made the container smaller, while keeping the temperature constant, what will happen?

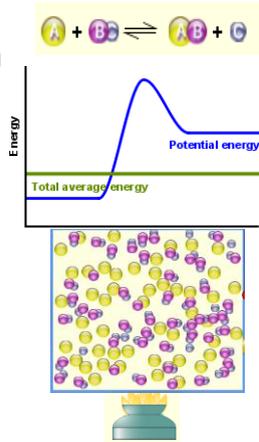


- The system will shift to the right
- The system will shift to the left
- This system would not effected



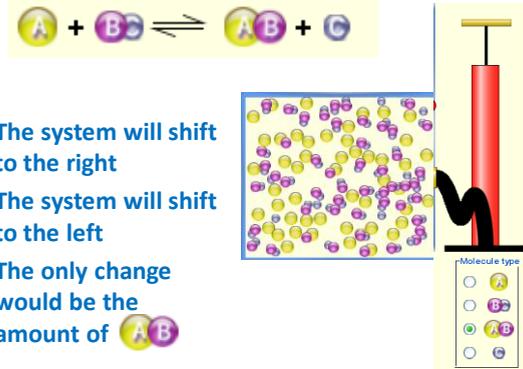
What would happen if you added energy using the heater ?

- The system will shift to the right
- The system will shift to the left
- Both reactants and products would have more energy, but the amounts would not change much



What would happen if you added AB ?

- The system will shift to the right
- The system will shift to the left
- The only change would be the amount of AB



What would happen if you added energy using the heater ?

- The system will shift to the right
- The system will shift to the left
- Both reactants and products would have more energy, but the amounts would not change much

