

# Reactions and Rates 3 Clicker Questions

## Activity 3: Introduction to Equilibrium

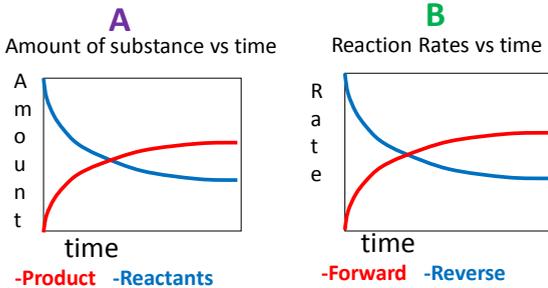
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PhET

### Learning Goals

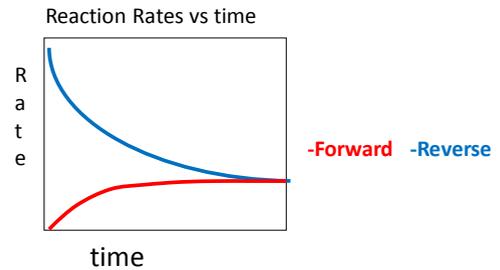
Students will be able to:

- Use a physical experiment to model chemical equilibrium
- Sketch how the number of reactants and products will change as a reaction proceeds
- Predict how changing the initial conditions will affect the equilibrium amounts of reactants and products.
- Predict how the shape of the reaction coordinate will affect the equilibrium amounts of reactants and products.

Which best shows that equilibrium has been reached?



Correct rate graph  
Forward reaction rate = Reverse rate



Which could show that equilibrium has been reached?

Select a reaction:  $A + B \rightleftharpoons C$

Start with how many...  
A? 50 BC? 50  
AB? 0 C? 0

Initial temperature: Cold Hot

End Experiment

Current Amounts:  
A: 27 B: 27 C: 23

**A**

Select a reaction:  $A + B \rightleftharpoons C$

Start with how many...  
A? 50 BC? 50  
AB? 0 C? 0

Initial temperature: Cold Hot

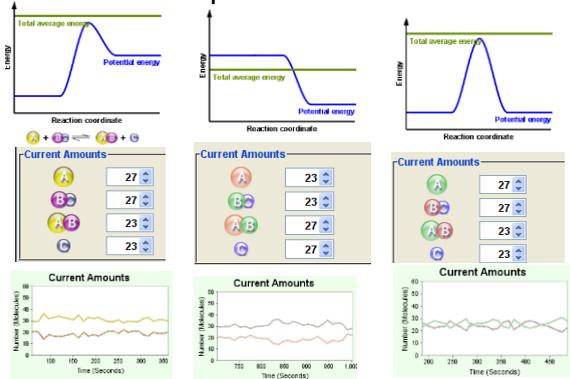
End Experiment

Current Amounts:  
A: 23 B: 23 C: 27

**B**

C neither  
D either

All are at equilibrium within limits

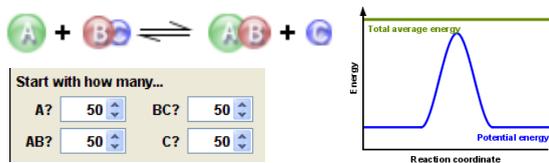
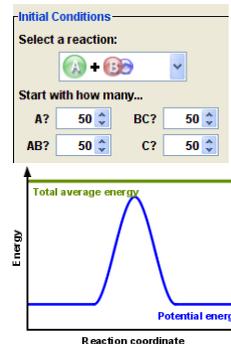


## Which best shows that equilibrium has been reached?

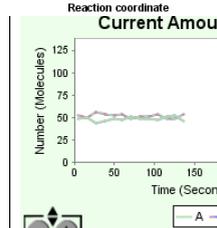
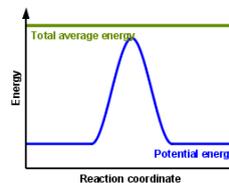
- The number of reactants is greater than the products
- The number of products is greater than the reactants
- The number of products is equal to the reactants
- The number of products varies little



At equilibrium, what would you predict is in the container?

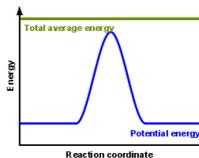


- Container will have mostly  $A$  &  $B$
- Container will have mostly  $AB$  &  $C$
- Container will have a mixture of all four with nearly equal amounts
- No reaction will occur since the products and reactants have the same energy

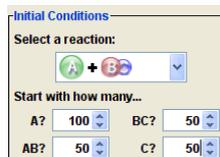
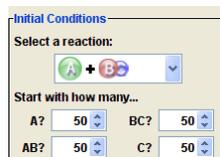


data

How will the equilibrium of second trial compare to the equilibrium of the first?



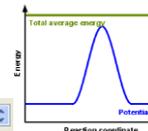
First experiment Second experiment



First trial



Second trial

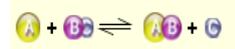
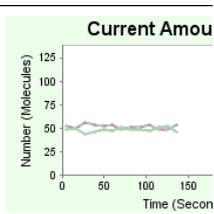


- There will be more  $A$  &  $B$
- There will be more  $AB$  &  $C$
- There will be more  $A$  &  $AB$  &  $C$
- There will be more  $B$  &  $AB$  &  $C$
- The ratios will still be about the same

Data for reactions

**Current Amounts**

A: 50  
 B: 50  
 AB: 50  
 C: 50



At equilibrium, what would you predict is in the container?

**Initial Conditions**

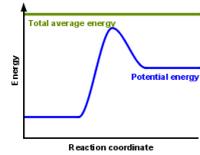
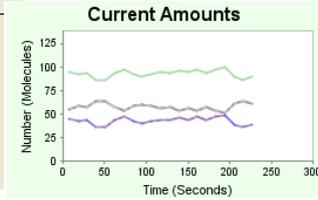
Select a reaction:  $A + B \rightleftharpoons AB + C$

Start with how many...

A? 100 BC? 100  
 AB? 0 C? 0

**Current Amounts**

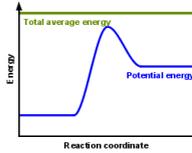
A: 92  
 B: 42  
 AB: 58  
 C: 58



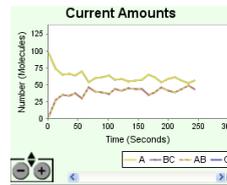
$A + B \rightleftharpoons AB + C$

Start with how many...

A? 100 BC? 100  
 AB? 0 C? 0



data



**Initial Conditions**

Select a reaction:  $A + B \rightleftharpoons AB + C$

Start with how many...

A? 100 BC? 100  
 AB? 0 C? 0

Initial temperature: Cold Hot

End Experiment

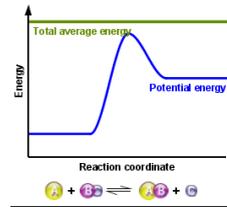
**Current Amounts**

A: 54  
 B: 54  
 AB: 46  
 C: 46

**Options**

Chart Options

Bar  Strip  
 Pie  None



- A. Container will have only  $AB$  &  $C$
- B. Container will have only  $A$  &  $B$
- C. Container will have a mixture of all four with more  $AB$  &  $C$
- D. Container will have a mixture of all four with more  $A$  &  $B$