

# Salts and Solubility Activity 1

**Learning Goals** Students will be able to:

- Determine the chemical formula by observation of ionic ratios in solutions
- Relate the simulation scale to real lab equipment through illustration and calculations
- Predict the chemical formula of compounds with a variety of ion charge combinations

Trish Loeblein July 2008 Questions 1-3 are a pretest. 4-8 are reflective

1. Which is the formula for the compound made from  $M^{+1}$  and  $N^{-2}$



2. Which is the formula for the compound made from

$M^{+3}$  and  $N^{-1}$



3. Which is the formula for the compound made from

$M^{+3}$  and  $N^{-2}$



4. I thought this lab was \_\_\_\_\_  
USEFUL for learning about ionic  
formulas.

A. very

B. mostly

C. barely

D. not

5. I thought this lab was \_\_\_\_\_  
ENJOYABLE for learning about ionic  
formulas.

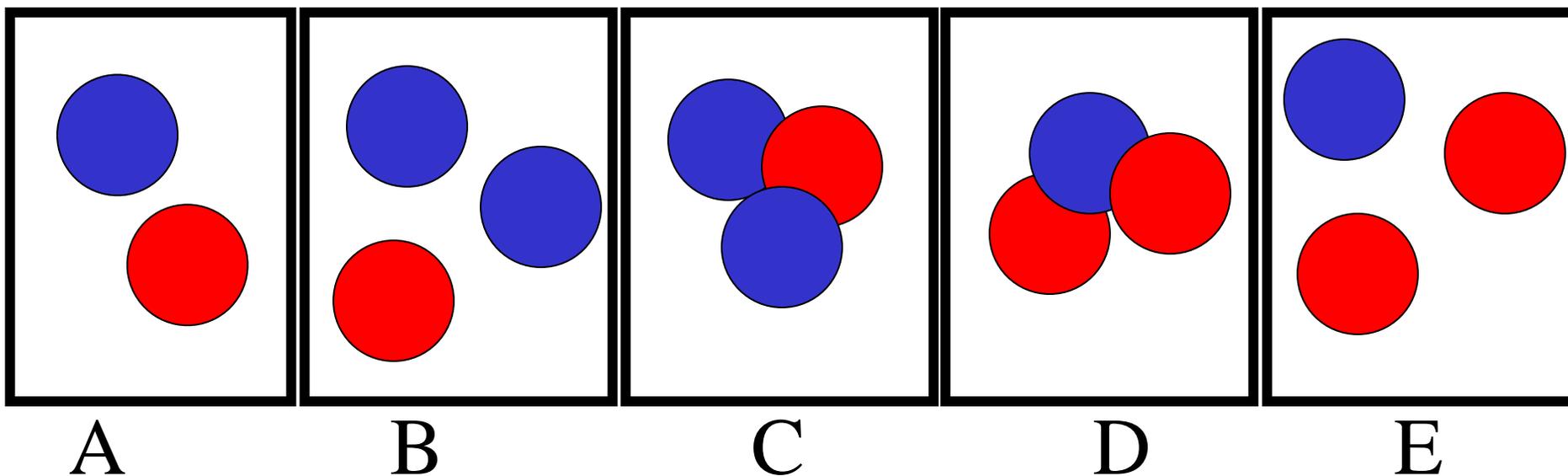
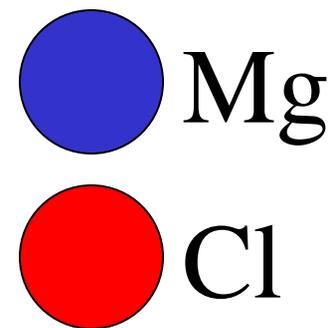
A. very

B. mostly

C. barely

D. not

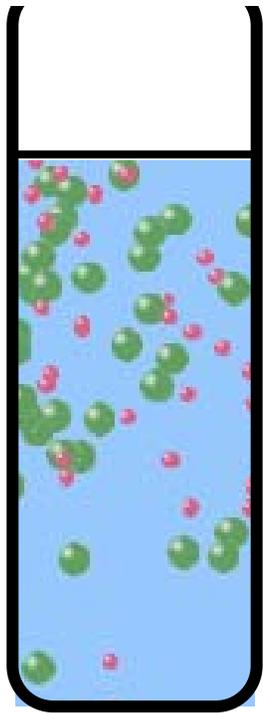
6. Which is the best drawing for Magnesium chloride in a water solution?



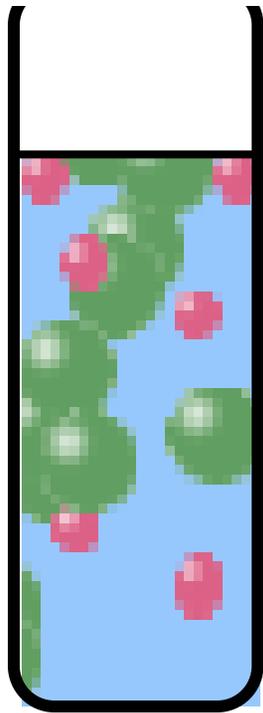
7. How would the drawing change if Magnesium chloride were changed to Magnesium oxide?

- A. The ratio of the ions would be the same
- B. The ratio would change to 1 magnesium for every oxide
- C. The ratio would change to 2 magnesium for every oxide
- D. You would have to use different colors

8. Which drawing best represents how large ions should be drawn in a 5 ml test tube of water?



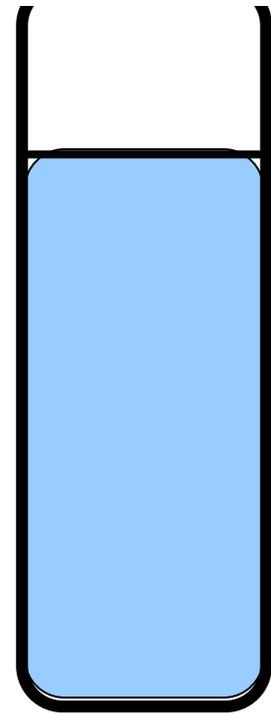
A



B



C



D